

		Anions								
		Oxide	_____ Cl <sup>-</sup>	_____	_____	_____	Phosphide	_____	_____ Se <sup>2-</sup>	Iodide
<b>Cations</b>	Iron (II)		FeCl <sub>2</sub>							
	Fe <sup>3+</sup>				Iron (III) Sulfide		Iron (III) Phosphide FeP			
		Calcium Oxide								
				SrBr <sub>2</sub>						
	Potassium						Potassium Nitride			KI
	Cu <sup>2+</sup>					CuF <sub>2</sub>		Cu <sub>3</sub> N <sub>2</sub>		
	Copper (I)		Copper (I) Chloride			CuF				
	Al <sup>3+</sup>								Aluminum Selenide	
	Lithium									
	Manganese (II)						Mn <sub>3</sub> P <sub>2</sub>			
	Mn <sup>7+</sup>		MnCl <sub>7</sub>							Manganese (VII) Iodide

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## Formulas of Ionic Compounds

Instructions: On the opposite side of this page you will find a large chart which shows how ionic compounds form from anions (negative ions) and cations (positive ion). Your task is to complete the chart with the appropriate information. Follow the steps below to help complete the chart:

- 1) The name of either the ion or compound should be written over each ion symbol or ionic formula unit. Some of the required information is already filled out.
- 2) All compounds are neutral, therefore the combined charge of the ions within the compound must add to zero.
- 3) The names of anions always end in -ide.
- 4) The stock system should be used if you see more than one type of cation. For instance Nickel has two different ions:  $\text{Ni}^{2+}$ , named Nickel (II) and  $\text{Ni}^{3+}$ , named Nickel (III). Use the octet rule to determine the charge for all other ions.
- 5) Have fun!