|  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- |
| 1 |  | 1. Write the element symbol and name in the bottom of the box.  2. Write the atomic mass number (rounded up to the nearest whole number) in the small upper right box for each element.  3. Draw a Bohr-Rutherford diagram with proton (P) and neutron (N) number in the middle of shells and fill the shells appropriately.   * Remember, two electrons in the first shell, 8 in the next shell, then 8 again and then 18. * Number of protons is atomic mass number – atomic number. | | | | | | 2 |
| 3  7  Li - Lithium | 4 |  | 5 | 6 | 7 | 8 | 9 | 10 |
| 11 | 12 |  | 13 | 14 | 15 | 16 | 17 | 18 |
| 19 | 20 |  |  |  |  |  | Halogen Family | Inert Gases Family |

The Periodic Table of Elements

Name:

Alkali Metals Alkaline Earth Metals

**The Periodic Table of Elements**

1. What is similar about the outer shell of electrons in the **Alkali Metals Family?**

2. What is similar about the outer shell of electrons in the **Alkaline Earth Metal Family?**

3. What is similar about the outer shell of electrons in the **Halogen Family?**

4. What is similar about the outer shell of electrons in the **Inert Gases Family? (also called**

**the Noble Gases)**

5. The rows are called periods. What is similar about the electron shells as you look across a

period?