# Types of Chemical Reactions for the second secon





Name	D	ate				
Types of Chemical Reactions Cut and Paste						
<b>Directions:</b> Cut the labels at the bottom of the page. P chemical reactions below.	lace the labels in the appro	priate space to identify the				
8 Fe+S <sub>8</sub> → 8 FeS						
$2 H_2O_2 \rightarrow 2 H_2O + O_2$						
Zn + 2 HCl → ZnCl <sub>2</sub> + H <sub>2</sub>						
$C_{10}H_8 + 12 O_2 \rightarrow 10 CO_2 + 4 H_2O_2$						
C <sub>10</sub> II <sub>8</sub> + 12 O <sub>2</sub> > 10 CO <sub>2</sub> + 4 II <sub>2</sub> O						
2 AgI + Na <sub>2</sub> S → Ag <sub>2</sub> S + 2 NaI						
$3 \operatorname{CaCl}_2 + 2 \operatorname{Na}_3 \operatorname{PO}_4 \rightarrow \operatorname{Ca}_3 (\operatorname{PO}_4)_2 + 6 \operatorname{NaCl}$						
$SnO_2 + 2H_2 \rightarrow Sn + 2H_2O$						
	Single	Double				

これ

and

Combustion	Synthesis	Single Replacement	Double Replacement
Combustion	Decomposition	Single Replacement	Double Replacement

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#### **Types of Chemical Reactions Cut and Paste**

*Directions:* Cut the labels at the bottom of the page. Place the labels in the appropriate space to identify the types of chemical reactions below.

8 Fe + S<sub>8</sub>  $\rightarrow$  8 FeS  $2 H_2O_2 \rightarrow 2 H_2O + O_2$  $Zn + 2 HCl \rightarrow ZnCl_2 + H_2$  $C_{10}H_8 + 12 O_2 \rightarrow 10 CO_2 + 4 H_2O_2$  $2 \text{ Agl} + \text{Na}_2\text{S} \rightarrow \text{Ag}_2\text{S} + 2 \text{ Nal}$  $3 \text{ CaCl}_2 + 2 \text{ Na}_3 \text{PO}_4 \rightarrow \text{Ca}_3 (\text{PO}_4)_2 + 6 \text{ NaCl}$  $SnO_2 + 2 H_2 \rightarrow Sn + 2 H_2O$ 

Combustion	Synthesis	Single Replacement	Double Replacement
Combustion	Decomposition	Single Replacement	Double Replacement

Date

### Types of Chemical Reactions Cut and Paste Key

**Directions:** Cut the labels at the bottom of the page. Place the labels in the appropriate space to identify the types of chemical reactions below.

8 Fe + S <sub>8</sub> $\rightarrow$ 8 FeS			Synthesis		
$2 H_2O_2 \rightarrow 2 H_2O + O_2$			Decomposition		
Zn + 2 HCl → ZnCl <sub>2</sub> + H <sub>2</sub>		Single Replacement			
$C_{10}H_8$ + 12 $O_2$ → 10 $CO_2$ + 4 $H_2O$			Combustion		
2 AgI + Na <sub>2</sub> S → Ag <sub>2</sub> S + 2 NaI			Double Replacement		
3 CaCl <sub>2</sub> + 2 Na <sub>3</sub> PO <sub>4</sub> → Ca <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> + 6 NaCl			Double Replacement		
$SnO_2 + 2 H_2 \rightarrow Sn + 2 H_2O$			Single Replacement		
r	r	<b></b>		·1	
Combustion	Synthesis	Re	Single Doub Replacement Replacer		
Combustion	Decomposition	Re	Single placement	Double Replacement	

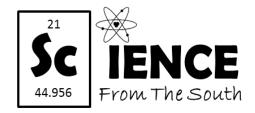
## Thank you for your download!

If you have any questions or concerns, please do not hesitate to contact me at <u>sciencefromthesouth@gmail.com</u>. Your 100% satisfaction is valued. Feedback and ratings are also greatly appreciated.

#### **Teacher's Notes:**

Use this download to pre-assess, review, or even assess your students' understanding of various types of chemical reactions.

- No prep required. Just print and go.
- Use it as a quick time filler or practice assignment.
- Use it as an assignment for early finishers.
- Use it as a pre-assessment to see what your students know about types of chemical reactions.
- Use it as an assessment, quiz, or test.
- Assign it as a more hands on homework assignment.
- Make it a part of a station activity or review.
- Answer key included for ease of checking. Use it to help you grade or print and give to students for self-assessing.



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